Effect of polymer-surface mobility on adhesion in poly(methyl methacrylate)-tape system

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effect of polymer-surface mobility on adhesion in poly (methyl acrylate)-tape system

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Introduction

The interaction between two polymer layers, especially adhesion between them, plays an important role in polymer processing and other applications. Detailed knowledge of the molecular structure and dynamics of polymer interfaces, and how they relate to macroscopic mechanical properties, should help designers construct more functional systems. Unfortunately, there have been few studies where both molecular and macroscopic studies have been performed on similar systems. In previous studies from our group, we have probed the dynamics1 of poly(methyl acrylate) (PMA) and thermal behavior of poly(methyl methacrylate) (PMMA) on silica. These studies helped us paint a picture for strongly bound molecules on silica, where a motional gradient perpendicular to the surface was observed. More mobile, lower Tg material was found at the air-polymer interface, while less mobile, higher Tg material was found at the polymer-substrate interface.

A previous study2 from our group showed that the glass-transition temperature of PMMA changes with the thickness of the polymer layer on silica. By examining the thermal behavior of the polymer with modulated differential scanning calorimetry (MDSC), it was observed that the glass-transition temperature increases as the thickness of the polymer layer decreases. Blum and Lin3 have also used the deuterium NMR to probe the dynamics of bulk and silica-absorbed poly(methyl acrylate)-d5 (PMA-d5). It was found that, an increase in the absorbed amount of polymer increased the mobility of polymer at the air–polymer interface and a decrease in the absorption of polymer, decreased the polymer mobility. It is not known how this change in mobility relates to the mechanical properties of polymers, especially adhesion between two polymer layers.

Of the various methods available for measuring adhesion-related properties, peeling is considered to be the most convenient.4 Current theories consider peeling to be the principal mode of separation of an adhesive from the substrate. It is proposed that adhesion strength is proportional to the surface free energy.5,6 Research has been done to find the adhesion strength between films and substrates, but there are some questions which remain unanswered like: Does adhesion strength depend on the thickness of the polymer layer? Does adhesion strength relate to the mobility of polymer segments on the surface?

In the present report, we summarize some of our macroscopic studies, namely peel tests, on systems similar to those previously studied. We find that the mobility of the polymer chains at the interface play an important role in determining the peel strength between the two polymer layers. We decided to use a 90° peel angle method. At this high angle, the lateral stress in the film is very negligible, thereby reducing the lateral elongation of the film. In addition, at this angle, there is no slip at the interface. Our studies show that as the polymer film thickness decreases, the force required to initiate fracture between two polymer surfaces increases and the fracture energy also increases proportionately. We have also observed how fracture energy changes with a change in film thickness.

Experimental

Poly(methyl acrylate) PMA was synthesized via bulk radical solution polymerization of methyl acrylate (Aldrich) with AIBN (recrystallized before use with methanol) as an initiator. Methyl acrylate was stirred overnight with CaH2 (calium hydride) and then vacuum distilled. 25 ml of methyl acrylate was mixed with 0.03 gm of AIBN and reaction was carried out at 60°C for 5 h. The properties measured for PMA are given in Table 1. The polymer was dissolved in toluene and precipitated with methanol. Glass slides (Corning Micro Slides 2948, single frosted) were used as the substrate for the experiments. PMMA solutions in toluene were prepared to coat the glass slides.

Table 1. Characterization of PMA

<table>
<thead>
<tr>
<th>Polymer</th>
<th>Tg °C</th>
<th>Mw (g/mol)</th>
<th>Polydispersity</th>
</tr>
</thead>
<tbody>
<tr>
<td>PMA</td>
<td>10</td>
<td>342,000</td>
<td>2.01</td>
</tr>
</tbody>
</table>

Results and Discussion

Peel force vs. peel velocity data for various film thicknesses are tabulated in Table 2. From the data, it can be seen that peel velocities increase as the amount of force used in separating two surfaces increases. Figure 1 shows the trend of peel force versus the peel velocity. A linear relationship between peel force and peel velocity is observed. The x intercepts give the amount of force required for initial crack formation in the system. At higher forces, the excess energy is used for the propagation of crack.

Figure 1. Force (N) vs velocity (µm/s) for various film thicknesses. The experiments were conducted at a 90° peel angle.

It was observed that the force required to separate the two polymer layers increases as the film thickness decreases. We believe that this phenomenon is related to the polymer mobility at the polymer-polymer interface and at the polymer–substrate interface. It is observed that in the vicinity of an interface, the dynamics of a polymer chain may be altered. The mobility of polymer is greater at the air-polymer interface because of the free volume available at that interface. In addition, for strongly bound polymers, such as those observed here,1 the polymer mobility is decreased at the substrate-polymer interface, where polymer sticks to the surface. Increasing the amount of polymer coated on the substrate, i.e. changing from thin films to thick films, the polymer mobility increases at the air-polymer interface. Blum and Lin3 studied the mobility of polymer segments at the polymer-polymer interface using deuterium NMR. These results show that mobility of polymer at the polymer-polymer interface decreases compared to the air-polymer interface because of the interaction between the two polymer layers. At the polymer-polymer interface, the mobility of polymer attached to the substrate

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Fractures can be of the cohesive or adhesive type. To ensure that the fracture is taking place at the interface (adhesive fracture), AFM studies of the samples were conducted before and after the peeling experiments. Tapping mode AFM was used to study the surface of the polymer. Figure 2 (a) shows PMA coated on the glass plate initially (before the peeling experiment) and Figure 2 (b) shows the PMA surface after the peeling experiment. No significant change in thickness of polymer layers was observed. No chain pull out appeared to occur.

Figure 3 shows the film thicknesses vs. fracture energy plot for different velocities. It is observed that the graph has the same trend for all the film thicknesses. The data is fit to power law dependence. The present treatment does not allow for a complete discussion of this fitting.

Conclusions

From the experimental data, we conclude that fracture energy depends on the film thickness. The mobility of the polymer is also a function of film thickness, and we believe that this means that the fracture energy depends on the mobility of polymer, since the amount of energy required to initiate fracture between two polymer layers increases as the polymer thickness decreases. The fracture energy also behaves in a manner similar to the peel force with changes in thickness. We have worked with PMA (rubbery polymer) and further research is likely to be with PMMA (glassy polymer).

References

1. Lin, W. Y.; and Blum, F. D. Macromolecules 1997, 30, 5331.

Fracture energy is defined as the energy applied to the system per unit extension of the crack area when fracture takes place in the system. The fracture energy was calculated using the Griffith energy balance criterion. According to the Griffith energy balance the fracture energy is a rate dependent quantity. Fracture energies were calculated using equation 1. In equation 1, \( P \) is peel force per unit width of film, \( E \) is Young’s modulus of film and \( h \) is the thickness of the film. The first term in equation 1 gives the amount of recoverable strain energy. For high peel angles (e.g. 90°) the first term of the equation represents the elastic strain energy, and so for our case the fracture energy is given by eq. 2.

\[
G_c = (\pi r_0^2 / 2Eh) + P (1 - \cos \theta) \\
G_c = P (1 - \cos \theta)
\]  

(1) 

(2)

In our case, the peel angle is 90°. So the fracture energy \( (G_c) \) is equal to the peel force per unit width of film \( (P) \). The fracture energy for different peel velocities is tabulated in Table 2. From the data, it is observed that the fracture energy also increases as the film thickness decreases. The intercept of peel velocities is tabulated in Table 2. From the data, it is observed that the graph has the same trend for all the film thicknesses. The data is fit to power law dependence.